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Phys 221 (Section 6)

Quiz #6

1. A trumpet player in an open automobile plays a note of frequency 392 Hz on his horn. A stationary observer directly ahead of the auto hears a note of frequency 440 Hz. What is the speed of the automobile?

If frequency f is emitted by a source moving toward a stationary observer then the latter hears a frequency

$$f' = \left(\frac{v}{v - v_s} \right) f \quad \text{where } v_s \text{ is the speed of the source, } v \text{ is the speed of sound.}$$

Solving for v_s ,

$$\left(\frac{v}{v - v_s} \right) = \frac{f'}{f} = \frac{440}{392} = 1.12$$

$$v = 1.12 (v - v_s)$$

$$1.12 v_s = 0.122 v$$

$$\rightarrow v_s = 0.109 v = 0.109 (343 \frac{m}{s}) = \boxed{37.4 \frac{m}{s}}$$



2. What mass of He (Helium) gas is contained in a volume of 0.3 m^3 at a pressure of 2.0 atm and a temperature of 20.0°C ? (Take He to be an ideal gas.)

$$P = 2.0 \text{ atm}, \quad V = 0.3 \text{ m}^3 = 0.3 \times 10^3 \text{ L}, \quad T = (20 + 273.15) \text{ K} = 293 \text{ K}$$

From the ideal gas law, the number of moles of He contained is

$$n = \frac{PV}{RT} = \frac{(2.0 \text{ atm})(0.3 \times 10^3 \text{ L})}{(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}})(293 \text{ K})} = 24.9 \text{ moles}$$

The mass of the amount of helium is

$$m = nM = (24.9 \text{ moles}) \left(\frac{4.0 \text{ g}}{\text{mol}} \right) = \boxed{99.8 \text{ g}} = 0.0998 \text{ kg}$$

3. How much heat is required to convert 3.0 g of ice at 0.0°C to water at 25.0°C ?

Heat required to change Ice, $0^\circ\text{C} \rightarrow$ Water, 0°C

$$Q_1 = mL_{fus} = (3.0 \times 10^{-3} \text{ kg}) (3.33 \times 10^5 \frac{\text{J}}{\text{kg}}) = 999 \text{ J}$$

Heat required to change Water, $0^\circ\text{C} \rightarrow$ Water, 25°C

$$Q_2 = mc \Delta T = (3.0 \times 10^{-3} \text{ kg}) (4186 \frac{\text{J}}{\text{kg}\cdot\text{C}^\circ}) (25 \text{ C}^\circ) = 314 \text{ J}$$

Total is

$$Q = Q_1 + Q_2 = \boxed{1313 \text{ J}}$$

$$\text{Speed of sound} = 343 \frac{\text{m}}{\text{s}}$$

$$R = 0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$$

$$N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$$

$$\text{Helium: } M = 4.0 \frac{\text{g}}{\text{mol}}$$

$$\text{Ice: } c = 2090 \frac{\text{J}}{\text{kg}\cdot\text{C}^\circ}$$

$$\text{Water: } c = 4186 \frac{\text{J}}{\text{kg}\cdot\text{C}^\circ}$$

$$\text{Water (H}_2\text{O): } L_{fus} = 3.33 \times 10^5 \frac{\text{J}}{\text{kg}}$$

$$L_{vap} = 2.26 \times 10^6 \frac{\text{J}}{\text{kg}}$$

$$T_C = T - 273.15$$

$$1.0 \text{ L} = 10^3 \text{ cm}^3$$